

## LAMPIRAN

Lampiran 1. Tabel A-1 *Properties Tables and Charts (SI units) Appendix 1*

**TABLE A – 1**

Molar mass, gas constant, and critical-point properties

Substance	Formula	Molar mass, <i>M</i> kg/kmol	Gas constant, <i>R</i> kJ/kg · K*	Critical-point properties		
				Temperature, K	Pressure, MPa	Volume, m <sup>3</sup> /kmol
Air	—	28.97	0.2870	132.5	3.77	0.0883
Ammonia	NH <sub>3</sub>	17.03	0.4882	405.5	11.28	0.0724
Argon	Ar	39.948	0.2081	151	4.86	0.0749
Benzene	C <sub>6</sub> H <sub>6</sub>	78.115	0.1064	562	4.92	0.2603
Bromine	Br <sub>2</sub>	159.808	0.0520	584	10.34	0.1355
<i>n</i> -Butane	C <sub>4</sub> H <sub>10</sub>	58.124	0.1430	425.2	3.80	0.2547
Carbon dioxide	CO <sub>2</sub>	44.01	0.1889	304.2	7.39	0.0943
Carbon monoxide	CO	28.011	0.2968	133	3.50	0.0930
Carbon tetrachloride	CCl <sub>4</sub>	153.82	0.05405	556.4	4.56	0.2759
Chlorine	Cl <sub>2</sub>	70.906	0.1173	417	7.71	0.1242
Chloroform	CHCl <sub>3</sub>	119.38	0.06964	536.6	5.47	0.2403
Dichlorodifluoromethane (R-12)	CCl <sub>2</sub> F <sub>2</sub>	120.91	0.06876	384.7	4.01	0.2179
Dichlorofluoromethane (R-21)	CHCl <sub>2</sub> F	102.92	0.08078	451.7	5.17	0.1973
Ethane	C <sub>2</sub> H <sub>6</sub>	30.070	0.2765	305.5	4.48	0.1480
Ethyl alcohol	C <sub>2</sub> H <sub>5</sub> OH	46.07	0.1805	516	6.38	0.1673
Ethylene	C <sub>2</sub> H <sub>4</sub>	28.054	0.2964	282.4	5.12	0.1242
Helium	He	4.003	2.0769	5.3	0.23	0.0578
<i>n</i> -Hexane	C <sub>6</sub> H <sub>14</sub>	86.179	0.09647	507.9	3.03	0.3677
Hydrogen (normal)	H <sub>2</sub>	2.016	4.1240	33.3	1.30	0.0649
Krypton	Kr	83.80	0.09921	209.4	5.50	0.0924
Methane	CH <sub>4</sub>	16.043	0.5182	191.1	4.64	0.0993
Methyl alcohol	CH <sub>3</sub> OH	32.042	0.2595	513.2	7.95	0.1180
Methyl chloride	CH <sub>3</sub> Cl	50.488	0.1647	416.3	6.68	0.1430
Neon	Ne	20.183	0.4119	44.5	2.73	0.0417
Nitrogen	N <sub>2</sub>	28.013	0.2968	126.2	3.39	0.0899
Nitrous oxide	N <sub>2</sub> O	44.013	0.1889	309.7	7.27	0.0961
Oxygen	O <sub>2</sub>	31.999	0.2598	154.8	5.08	0.0780
Propane	C <sub>3</sub> H <sub>8</sub>	44.097	0.1885	370	4.26	0.1998
Propylene	C <sub>3</sub> H <sub>6</sub>	42.081	0.1976	365	4.62	0.1810
Sulfur dioxide	SO <sub>2</sub>	64.063	0.1298	430.7	7.88	0.1217
Tetrafluoroethane (R-134a)	CF <sub>3</sub> CH <sub>2</sub> F	102.03	0.08149	374.2	4.059	0.1993
Trichlorofluoromethane (R-11)	CCl <sub>3</sub> F	137.37	0.06052	471.2	4.38	0.2478
Water	H <sub>2</sub> O	18.015	0.4615	647.1	22.06	0.0560
Xenon	Xe	131.30	0.06332	289.8	5.88	0.1186

\*The unit kJ/kg · K is equivalent to kPa · m<sup>3</sup>/kg · K. The gas constant is calculated from  $R = R_u/M$ , where  $R_u = 8.31447$  kJ/kmol · K and  $M$  is the molar mass.

Source: K. A. Kobe and R. E. Lynn, Jr., *Chemical Review* 52 (1953), pp. 117–236; and ASHRAE, *Handbook of Fundamentals* (Atlanta, GA: American Society of Heating, Refrigerating and Air-Conditioning Engineers, Inc., 1993), pp. 16.4 and 36.1.

Lampiran 2. Tabel A-2a *Properties Tables and Charts (SI units) Appendix 1*

**TABLE A-2**

Ideal-gas specific heats of various common gases

(a) At 300 K

Gas	Formula	Gas constant, $R$ kJ/kg · K	$c_p$ kJ/kg · K	$c_v$ kJ/kg · K	$k$
Air	—	0.2870	1.005	0.718	1.400
Argon	Ar	0.2081	0.5203	0.3122	1.667
Butane	C <sub>4</sub> H <sub>10</sub>	0.1433	1.7164	1.5734	1.091
Carbon dioxide	CO <sub>2</sub>	0.1889	0.846	0.657	1.289
Carbon monoxide	CO	0.2968	1.040	0.744	1.400
Ethane	C <sub>2</sub> H <sub>6</sub>	0.2765	1.7662	1.4897	1.186
Ethylene	C <sub>2</sub> H <sub>4</sub>	0.2964	1.5482	1.2518	1.237
Helium	He	2.0769	5.1926	3.1156	1.667
Hydrogen	H <sub>2</sub>	4.1240	14.307	10.183	1.405
Methane	CH <sub>4</sub>	0.5182	2.2537	1.7354	1.299
Neon	Ne	0.4119	1.0299	0.6179	1.667
Nitrogen	N <sub>2</sub>	0.2968	1.039	0.743	1.400
Octane	C <sub>8</sub> H <sub>18</sub>	0.0729	1.7113	1.6385	1.044
Oxygen	O <sub>2</sub>	0.2598	0.918	0.658	1.395
Propane	C <sub>3</sub> H <sub>8</sub>	0.1885	1.6794	1.4909	1.126
Steam	H <sub>2</sub> O	0.4615	1.8723	1.4108	1.327

Note: The unit kJ/kg · K is equivalent to kJ/kg · °C.

Source: *Chemical and Process Thermodynamics 3/E* by Kyle, B. G., © 2000. Adapted by permission of Pearson Education, Inc., Upper Saddle River, NJ.

Lampiran 3. Tabel A-2b *Properties Tables and Charts (SI units) Appendix 1*

**TABLE A-2**

Ideal-gas specific heats of various common gases (*Continued*)

(b) At various temperatures

Temperature, K	$c_p$	$c_v$	$k$	$c_p$	$c_v$	$k$	$c_p$	$c_v$	$k$
	kJ/kg · K	kJ/kg · K		kJ/kg · K	kJ/kg · K		kJ/kg · K		
	<i>Air</i>			<i>Carbon dioxide, CO<sub>2</sub></i>			<i>Carbon monoxide, CO</i>		
250	1.003	0.716	1.401	0.791	0.602	1.314	1.039	0.743	1.400
300	1.005	0.718	1.400	0.846	0.657	1.288	1.040	0.744	1.399
350	1.008	0.721	1.398	0.895	0.706	1.268	1.043	0.746	1.398
400	1.013	0.726	1.395	0.939	0.750	1.252	1.047	0.751	1.395
450	1.020	0.733	1.391	0.978	0.790	1.239	1.054	0.757	1.392
500	1.029	0.742	1.387	1.014	0.825	1.229	1.063	0.767	1.387
550	1.040	0.753	1.381	1.046	0.857	1.220	1.075	0.778	1.382
600	1.051	0.764	1.376	1.075	0.886	1.213	1.087	0.790	1.376
650	1.063	0.776	1.370	1.102	0.913	1.207	1.100	0.803	1.370
700	1.075	0.788	1.364	1.126	0.937	1.202	1.113	0.816	1.364
750	1.087	0.800	1.359	1.148	0.959	1.197	1.126	0.829	1.358
800	1.099	0.812	1.354	1.169	0.980	1.193	1.139	0.842	1.353
900	1.121	0.834	1.344	1.204	1.015	1.186	1.163	0.866	1.343
1000	1.142	0.855	1.336	1.234	1.045	1.181	1.185	0.888	1.335
	<i>Hydrogen, H<sub>2</sub></i>			<i>Nitrogen, N<sub>2</sub></i>			<i>Oxygen, O<sub>2</sub></i>		
250	14.051	9.927	1.416	1.039	0.742	1.400	0.913	0.653	1.398
300	14.307	10.183	1.405	1.039	0.743	1.400	0.918	0.658	1.395
350	14.427	10.302	1.400	1.041	0.744	1.399	0.928	0.668	1.389
400	14.476	10.352	1.398	1.044	0.747	1.397	0.941	0.681	1.382
450	14.501	10.377	1.398	1.049	0.752	1.395	0.956	0.696	1.373
500	14.513	10.389	1.397	1.056	0.759	1.391	0.972	0.712	1.365
550	14.530	10.405	1.396	1.065	0.768	1.387	0.988	0.728	1.358
600	14.546	10.422	1.396	1.075	0.778	1.382	1.003	0.743	1.350
650	14.571	10.447	1.395	1.086	0.789	1.376	1.017	0.758	1.343
700	14.604	10.480	1.394	1.098	0.801	1.371	1.031	0.771	1.337
750	14.645	10.521	1.392	1.110	0.813	1.365	1.043	0.783	1.332
800	14.695	10.570	1.390	1.121	0.825	1.360	1.054	0.794	1.327
900	14.822	10.698	1.385	1.145	0.849	1.349	1.074	0.814	1.319
1000	14.983	10.859	1.380	1.167	0.870	1.341	1.090	0.830	1.313

Source: Kenneth Wark, *Thermodynamics*, 4th ed. (New York: McGraw-Hill, 1983), p. 783, Table A-4M. Originally published in *Tables of Thermal Properties of Gases*, NBS Circular 564, 1955.

Lampiran 4. Tabel A-2c *Properties Tables and Charts (SI units) Appendix 1*

**TABLE A-2**

Ideal-gas specific heats of various common gases (*Concluded*)

(c) As a function of temperature

$$\bar{c}_p = a + bT + cT^2 + dT^3$$

( $T$  in K,  $c_p$  in kJ/kmol · K)

Substance	Formula	$a$	$b$	$c$	$d$	Temperature range, K	% error	
							Max.	Avg.
Nitrogen	N <sub>2</sub>	28.90	-0.1571 × 10 <sup>-2</sup>	0.8081 × 10 <sup>-5</sup>	-2.873 × 10 <sup>-9</sup>	273–1800	0.59	0.34
Oxygen	O <sub>2</sub>	25.48	1.520 × 10 <sup>-2</sup>	-0.7155 × 10 <sup>-5</sup>	1.312 × 10 <sup>-9</sup>	273–1800	1.19	0.28
Air	—	28.11	0.1967 × 10 <sup>-2</sup>	0.4802 × 10 <sup>-5</sup>	-1.966 × 10 <sup>-9</sup>	273–1800	0.72	0.33
Hydrogen	H <sub>2</sub>	29.11	-0.1916 × 10 <sup>-2</sup>	0.4003 × 10 <sup>-5</sup>	-0.8704 × 10 <sup>-9</sup>	273–1800	1.01	0.26
Carbon monoxide	CO	28.16	0.1675 × 10 <sup>-2</sup>	0.5372 × 10 <sup>-5</sup>	-2.222 × 10 <sup>-9</sup>	273–1800	0.89	0.37
Carbon dioxide	CO <sub>2</sub>	22.26	5.981 × 10 <sup>-2</sup>	-3.501 × 10 <sup>-5</sup>	7.469 × 10 <sup>-9</sup>	273–1800	0.67	0.22
Water vapor	H <sub>2</sub> O	32.24	0.1923 × 10 <sup>-2</sup>	1.055 × 10 <sup>-5</sup>	-3.595 × 10 <sup>-9</sup>	273–1800	0.53	0.24
Nitric oxide	NO	29.34	-0.09395 × 10 <sup>-2</sup>	0.9747 × 10 <sup>-5</sup>	-4.187 × 10 <sup>-9</sup>	273–1500	0.97	0.36
Nitrous oxide	N <sub>2</sub> O	24.11	5.8632 × 10 <sup>-2</sup>	-3.562 × 10 <sup>-5</sup>	10.58 × 10 <sup>-9</sup>	273–1500	0.59	0.26
Nitrogen dioxide	NO <sub>2</sub>	22.9	5.715 × 10 <sup>-2</sup>	-3.52 × 10 <sup>-5</sup>	7.87 × 10 <sup>-9</sup>	273–1500	0.46	0.18
Ammonia	NH <sub>3</sub>	27.568	2.5630 × 10 <sup>-2</sup>	0.99072 × 10 <sup>-5</sup>	-6.6909 × 10 <sup>-9</sup>	273–1500	0.91	0.36
Sulfur	S <sub>2</sub>	27.21	2.218 × 10 <sup>-2</sup>	-1.628 × 10 <sup>-5</sup>	3.986 × 10 <sup>-9</sup>	273–1800	0.99	0.38
Sulfur dioxide	SO <sub>2</sub>	25.78	5.795 × 10 <sup>-2</sup>	-3.812 × 10 <sup>-5</sup>	8.612 × 10 <sup>-9</sup>	273–1800	0.45	0.24
Sulfur trioxide	SO <sub>3</sub>	16.40	14.58 × 10 <sup>-2</sup>	-11.20 × 10 <sup>-5</sup>	32.42 × 10 <sup>-9</sup>	273–1300	0.29	0.13
Acetylene	C <sub>2</sub> H <sub>2</sub>	21.8	9.2143 × 10 <sup>-2</sup>	-6.527 × 10 <sup>-5</sup>	18.21 × 10 <sup>-9</sup>	273–1500	1.46	0.59
Benzene	C <sub>6</sub> H <sub>6</sub>	-36.22	48.475 × 10 <sup>-2</sup>	-31.57 × 10 <sup>-5</sup>	77.62 × 10 <sup>-9</sup>	273–1500	0.34	0.20
Methanol	CH <sub>3</sub> O	19.0	9.152 × 10 <sup>-2</sup>	-1.22 × 10 <sup>-5</sup>	-8.039 × 10 <sup>-9</sup>	273–1000	0.18	0.08
Ethanol	C <sub>2</sub> H <sub>5</sub> O	19.9	20.96 × 10 <sup>-2</sup>	-10.38 × 10 <sup>-5</sup>	20.05 × 10 <sup>-9</sup>	273–1500	0.40	0.22
Hydrogen chloride	HCl	30.33	-0.7620 × 10 <sup>-2</sup>	1.327 × 10 <sup>-5</sup>	-4.338 × 10 <sup>-9</sup>	273–1500	0.22	0.08
Methane	CH <sub>4</sub>	19.89	5.024 × 10 <sup>-2</sup>	1.269 × 10 <sup>-5</sup>	-11.01 × 10 <sup>-9</sup>	273–1500	1.33	0.57
Ethane	C <sub>2</sub> H <sub>6</sub>	6.900	17.27 × 10 <sup>-2</sup>	-6.406 × 10 <sup>-5</sup>	7.285 × 10 <sup>-9</sup>	273–1500	0.83	0.28
Propane	C <sub>3</sub> H <sub>8</sub>	-4.04	30.48 × 10 <sup>-2</sup>	-15.72 × 10 <sup>-5</sup>	31.74 × 10 <sup>-9</sup>	273–1500	0.40	0.12
<i>n</i> -Butane	C <sub>4</sub> H <sub>10</sub>	3.96	37.15 × 10 <sup>-2</sup>	-18.34 × 10 <sup>-5</sup>	35.00 × 10 <sup>-9</sup>	273–1500	0.54	0.24
<i>i</i> -Butane	C <sub>4</sub> H <sub>10</sub>	-7.913	41.60 × 10 <sup>-2</sup>	-23.01 × 10 <sup>-5</sup>	49.91 × 10 <sup>-9</sup>	273–1500	0.25	0.13
<i>n</i> -Pentane	C <sub>5</sub> H <sub>12</sub>	6.774	45.43 × 10 <sup>-2</sup>	-22.46 × 10 <sup>-5</sup>	42.29 × 10 <sup>-9</sup>	273–1500	0.56	0.21
<i>n</i> -Hexane	C <sub>6</sub> H <sub>14</sub>	6.938	55.22 × 10 <sup>-2</sup>	-28.65 × 10 <sup>-5</sup>	57.69 × 10 <sup>-9</sup>	273–1500	0.72	0.20
Ethylene	C <sub>2</sub> H <sub>4</sub>	3.95	15.64 × 10 <sup>-2</sup>	-8.344 × 10 <sup>-5</sup>	17.67 × 10 <sup>-9</sup>	273–1500	0.54	0.13
Propylene	C <sub>3</sub> H <sub>6</sub>	3.15	23.83 × 10 <sup>-2</sup>	-12.18 × 10 <sup>-5</sup>	24.62 × 10 <sup>-9</sup>	273–1500	0.73	0.17

Source: B. G. Kyle, *Chemical and Process Thermodynamics* (Englewood Cliffs, NJ: Prentice-Hall, 1984). Used with permission.

Lampiran 5. Tabel A-3a *Properties Tables and Charts (SI units) Appendix 1*

TABLE A-3							
Properties of common liquids, solids, and foods							
(a) Liquids							
Substance	Boiling data at 1 atm		Freezing data		Liquid properties		
	Normal boiling point, °C	Latent heat of vaporization $h_{fg}$ , kJ/kg	Freezing point, °C	Latent heat of fusion $h_{if}$ , kJ/kg	Temperature, °C	Density $\rho$ , kg/m <sup>3</sup>	Specific heat $c_p$ , kJ/kg · K
Ammonia	-33.3	1357	-77.7	322.4	-33.3 -20 0 25	682 665 639 602	4.43 4.52 4.60 4.80
Argon	-185.9	161.6	-189.3	28	-185.6	1394	1.14
Benzene	80.2	394	5.5	126	20	879	1.72
Brine (20% sodium chloride by mass)	103.9	—	-17.4	—	20	1150	3.11
<i>n</i> -Butane	-0.5	385.2	-138.5	80.3	-0.5	601	2.31
Carbon dioxide	-78.4*	230.5 (at 0°C)	-56.6	—	0	298	0.59
Ethanol	78.2	838.3	-114.2	109	25	783	2.46
Ethyl alcohol	78.6	855	-156	108	20	789	2.84
Ethylene glycol	198.1	800.1	-10.8	181.1	20	1109	2.84
Glycerine	179.9	974	18.9	200.6	20	1261	2.32
Helium	-268.9	22.8	—	—	-268.9	146.2	22.8
Hydrogen	-252.8	445.7	-259.2	59.5	-252.8	70.7	10.0
Isobutane	-11.7	367.1	-160	105.7	-11.7	593.8	2.28
Kerosene	204–293	251	-24.9	—	20	820	2.00
Mercury	356.7	294.7	-38.9	11.4	25	13,560	0.139
Methane	-161.5	510.4	-182.2	58.4	-161.5 -100	423 301	3.49 5.79
Methanol	64.5	1100	-97.7	99.2	25	787	2.55
Nitrogen	-195.8	198.6	-210	25.3	-195.8 -160	809 596	2.06 2.97
Octane	124.8	306.3	-57.5	180.7	20	703	2.10
Oil (light)	—	—	—	—	25	910	1.80
Oxygen	-183	212.7	-218.8	13.7	-183	1141	1.71
Petroleum	—	230–384	—	—	20	640	2.0
Propane	-42.1	427.8	-187.7	80.0	-42.1 0 50	581 529 449	2.25 2.53 3.13
Refrigerant-134a	-26.1	217.0	-96.6	—	-50 -26.1 0	1443 1374 1295	1.23 1.27 1.34
Water	100	2257	0.0	333.7	25 0 25 50 75 100	997 1000 997 988 975 958	4.18 4.22 4.18 4.18 4.19 4.22

\* Sublimation temperature. (At pressures below the triple-point pressure of 518 kPa, carbon dioxide exists as a solid or gas. Also, the freezing-point temperature of carbon dioxide is the triple-point temperature of -56.5°C.)

Lampiran 6. Tabel A-15 *Properties Tables and Charts (SI units) Appendix 1*

**TABLE A-15**

Properties of air at 1 atm pressure

Temp. $T, ^\circ\text{C}$	Density $\rho, \text{kg/m}^3$	Specific Heat $c_p, \text{J/kg} \cdot \text{K}$	Thermal Conductivity $k, \text{W/m} \cdot \text{K}$	Thermal Diffusivity $\alpha, \text{m}^2/\text{s}^2$	Dynamic Viscosity $\mu, \text{kg/m} \cdot \text{s}$	Kinematic Viscosity $\nu, \text{m}^2/\text{s}$	Prandtl Number Pr
-150	2.866	983	0.01171	$4.158 \times 10^{-6}$	$8.636 \times 10^{-6}$	$3.013 \times 10^{-6}$	0.7246
-100	2.038	966	0.01582	$8.036 \times 10^{-6}$	$1.189 \times 10^{-5}$	$5.837 \times 10^{-6}$	0.7263
-50	1.582	999	0.01979	$1.252 \times 10^{-5}$	$1.474 \times 10^{-5}$	$9.319 \times 10^{-6}$	0.7440
-40	1.514	1002	0.02057	$1.356 \times 10^{-5}$	$1.527 \times 10^{-5}$	$1.008 \times 10^{-5}$	0.7436
-30	1.451	1004	0.02134	$1.465 \times 10^{-5}$	$1.579 \times 10^{-5}$	$1.087 \times 10^{-5}$	0.7425
-20	1.394	1005	0.02211	$1.578 \times 10^{-5}$	$1.630 \times 10^{-5}$	$1.169 \times 10^{-5}$	0.7408
-10	1.341	1006	0.02288	$1.696 \times 10^{-5}$	$1.680 \times 10^{-5}$	$1.252 \times 10^{-5}$	0.7387
0	1.292	1006	0.02364	$1.818 \times 10^{-5}$	$1.729 \times 10^{-5}$	$1.338 \times 10^{-5}$	0.7362
5	1.269	1006	0.02401	$1.880 \times 10^{-5}$	$1.754 \times 10^{-5}$	$1.382 \times 10^{-5}$	0.7350
10	1.246	1006	0.02439	$1.944 \times 10^{-5}$	$1.778 \times 10^{-5}$	$1.426 \times 10^{-5}$	0.7336
15	1.225	1007	0.02476	$2.009 \times 10^{-5}$	$1.802 \times 10^{-5}$	$1.470 \times 10^{-5}$	0.7323
20	1.204	1007	0.02514	$2.074 \times 10^{-5}$	$1.825 \times 10^{-5}$	$1.516 \times 10^{-5}$	0.7309
25	1.184	1007	0.02551	$2.141 \times 10^{-5}$	$1.849 \times 10^{-5}$	$1.562 \times 10^{-5}$	0.7296
30	1.164	1007	0.02588	$2.208 \times 10^{-5}$	$1.872 \times 10^{-5}$	$1.608 \times 10^{-5}$	0.7282
35	1.145	1007	0.02625	$2.277 \times 10^{-5}$	$1.895 \times 10^{-5}$	$1.655 \times 10^{-5}$	0.7268
40	1.127	1007	0.02662	$2.346 \times 10^{-5}$	$1.918 \times 10^{-5}$	$1.702 \times 10^{-5}$	0.7255
45	1.109	1007	0.02699	$2.416 \times 10^{-5}$	$1.941 \times 10^{-5}$	$1.750 \times 10^{-5}$	0.7241
50	1.092	1007	0.02735	$2.487 \times 10^{-5}$	$1.963 \times 10^{-5}$	$1.798 \times 10^{-5}$	0.7228
60	1.059	1007	0.02808	$2.632 \times 10^{-5}$	$2.008 \times 10^{-5}$	$1.896 \times 10^{-5}$	0.7202
70	1.028	1007	0.02881	$2.780 \times 10^{-5}$	$2.052 \times 10^{-5}$	$1.995 \times 10^{-5}$	0.7177
80	0.9994	1008	0.02953	$2.931 \times 10^{-5}$	$2.096 \times 10^{-5}$	$2.097 \times 10^{-5}$	0.7154
90	0.9718	1008	0.03024	$3.086 \times 10^{-5}$	$2.139 \times 10^{-5}$	$2.201 \times 10^{-5}$	0.7132
100	0.9458	1009	0.03095	$3.243 \times 10^{-5}$	$2.181 \times 10^{-5}$	$2.306 \times 10^{-5}$	0.7111
120	0.8977	1011	0.03235	$3.565 \times 10^{-5}$	$2.264 \times 10^{-5}$	$2.522 \times 10^{-5}$	0.7073
140	0.8542	1013	0.03374	$3.898 \times 10^{-5}$	$2.345 \times 10^{-5}$	$2.745 \times 10^{-5}$	0.7041
160	0.8148	1016	0.03511	$4.241 \times 10^{-5}$	$2.420 \times 10^{-5}$	$2.975 \times 10^{-5}$	0.7014
180	0.7788	1019	0.03646	$4.593 \times 10^{-5}$	$2.504 \times 10^{-5}$	$3.212 \times 10^{-5}$	0.6992
200	0.7459	1023	0.03779	$4.954 \times 10^{-5}$	$2.577 \times 10^{-5}$	$3.455 \times 10^{-5}$	0.6974
250	0.6746	1033	0.04104	$5.890 \times 10^{-5}$	$2.760 \times 10^{-5}$	$4.091 \times 10^{-5}$	0.6946
300	0.6158	1044	0.04418	$6.871 \times 10^{-5}$	$2.934 \times 10^{-5}$	$4.765 \times 10^{-5}$	0.6935
350	0.5664	1056	0.04721	$7.892 \times 10^{-5}$	$3.101 \times 10^{-5}$	$5.475 \times 10^{-5}$	0.6937
400	0.5243	1069	0.05015	$8.951 \times 10^{-5}$	$3.261 \times 10^{-5}$	$6.219 \times 10^{-5}$	0.6948
450	0.4880	1081	0.05298	$1.004 \times 10^{-4}$	$3.415 \times 10^{-5}$	$6.997 \times 10^{-5}$	0.6965
500	0.4565	1093	0.05572	$1.117 \times 10^{-4}$	$3.563 \times 10^{-5}$	$7.806 \times 10^{-5}$	0.6986
600	0.4042	1115	0.06093	$1.352 \times 10^{-4}$	$3.846 \times 10^{-5}$	$9.515 \times 10^{-5}$	0.7037
700	0.3627	1135	0.06581	$1.598 \times 10^{-4}$	$4.111 \times 10^{-5}$	$1.133 \times 10^{-4}$	0.7092
800	0.3289	1153	0.07037	$1.855 \times 10^{-4}$	$4.362 \times 10^{-5}$	$1.326 \times 10^{-4}$	0.7149
900	0.3008	1169	0.07465	$2.122 \times 10^{-4}$	$4.600 \times 10^{-5}$	$1.529 \times 10^{-4}$	0.7206
1000	0.2772	1184	0.07868	$2.398 \times 10^{-4}$	$4.826 \times 10^{-5}$	$1.741 \times 10^{-4}$	0.7260
1500	0.1990	1234	0.09599	$3.908 \times 10^{-4}$	$5.817 \times 10^{-5}$	$2.922 \times 10^{-4}$	0.7478
2000	0.1553	1264	0.11113	$5.664 \times 10^{-4}$	$6.630 \times 10^{-5}$	$4.270 \times 10^{-4}$	0.7539

Note: For ideal gases, the properties  $c_p$ ,  $k$ ,  $\mu$ , and Pr are independent of pressure. The properties  $\rho$ ,  $\nu$ , and  $\alpha$  at a pressure  $P$  (in atm) other than 1 atm are determined by multiplying the values of  $\rho$  at the given temperature by  $P$  and by dividing  $\nu$  and  $\alpha$  by  $P$ .

Source: Data generated from the EES software developed by S. A. Klein and F. L. Alvarado. Original sources: Keenan, Chao, Keyes, Gas Tables, Wiley, 198; and Thermophysical Properties of Matter, Vol. 3: Thermal Conductivity, Y. S. Touloukian, P. E. Liley, S. C. Saxena, Vol. 11: Viscosity, Y. S. Touloukian, S. C. Saxena, and P. Hestermans, IFI/Plenum, NY, 1970, ISBN 0-306067020-8.

Lampiran 7. Tabel A-26 *Properties Tables and Charts (SI units) Appendix 2*

**TABLE A-26E**

Enthalpy of formation, Gibbs function of formation, and absolute entropy at 77°F, 1 atm

Substance	Formula	$\bar{h}_f^\circ$ Btu/lbmol	$\bar{g}_f^\circ$ Btu/lbmol	$\bar{s}^\circ$ Btu/lbmol · R
Carbon	C(s)	0	0	1.36
Hydrogen	H <sub>2</sub> (g)	0	0	31.21
Nitrogen	N <sub>2</sub> (g)	0	0	45.77
Oxygen	O <sub>2</sub> (g)	0	0	49.00
Carbon monoxide	CO(g)	-47,540	-59,010	47.21
Carbon dioxide	CO <sub>2</sub> (g)	-169,300	-169,680	51.07
Water vapor	H <sub>2</sub> O(g)	-104,040	-98,350	45.11
Water	H <sub>2</sub> O(l)	-122,970	-102,040	16.71
Hydrogen peroxide	H <sub>2</sub> O <sub>2</sub> (g)	-58,640	-45,430	55.60
Ammonia	NH <sub>3</sub> (g)	-19,750	-7,140	45.97
Methane	CH <sub>4</sub> (g)	-32,210	-21,860	44.49
Acetylene	C <sub>2</sub> H <sub>2</sub> (g)	+97,540	+87,990	48.00
Ethylene	C <sub>2</sub> H <sub>4</sub> (g)	+22,490	+29,306	52.54
Ethane	C <sub>2</sub> H <sub>6</sub> (g)	-36,420	-14,150	54.85
Propylene	C <sub>3</sub> H <sub>6</sub> (g)	+8,790	+26,980	63.80
Propane	C <sub>3</sub> H <sub>8</sub> (g)	-44,680	-10,105	64.51
<i>n</i> -Butane	C <sub>4</sub> H <sub>10</sub> (g)	-54,270	-6,760	74.11
<i>n</i> -Octane	C <sub>8</sub> H <sub>18</sub> (g)	-89,680	+7,110	111.55
<i>n</i> -Octane	C <sub>8</sub> H <sub>18</sub> (l)	-107,530	+2,840	86.23
<i>n</i> -Dodecane	C <sub>12</sub> H <sub>26</sub> (g)	-125,190	+21,570	148.86
Benzene	C <sub>6</sub> H <sub>6</sub> (g)	+35,680	+55,780	64.34
Methyl alcohol	CH <sub>3</sub> OH(g)	-86,540	-69,700	57.29
Methyl alcohol	CH <sub>3</sub> OH(l)	-102,670	-71,570	30.30
Ethyl alcohol	C <sub>2</sub> H <sub>5</sub> OH(g)	-101,230	-72,520	67.54
Ethyl alcohol	C <sub>2</sub> H <sub>5</sub> OH(l)	-119,470	-75,240	38.40
Oxygen	O(g)	+107,210	+99,710	38.47
Hydrogen	H(g)	+93,780	+87,460	27.39
Nitrogen	N(g)	+203,340	+195,970	36.61
Hydroxyl	OH(g)	+16,790	+14,750	43.92

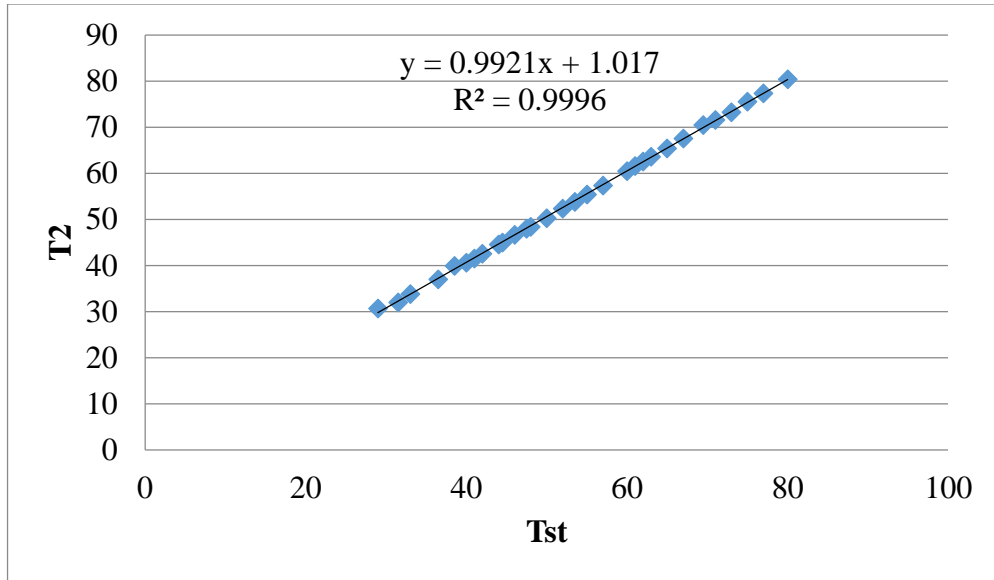
Source: From JANAF, *Thermochemical Tables* (Midland, MI: Dow Chemical Co., 1971), *Selected Values of Chemical Thermodynamic Properties*, NBS Technical Note 270-3, 1968; and *API Research Project 44* (Carnegie Press, 1953).

Lampiran 8. Tabel *Emissivities of some material at 300 K*

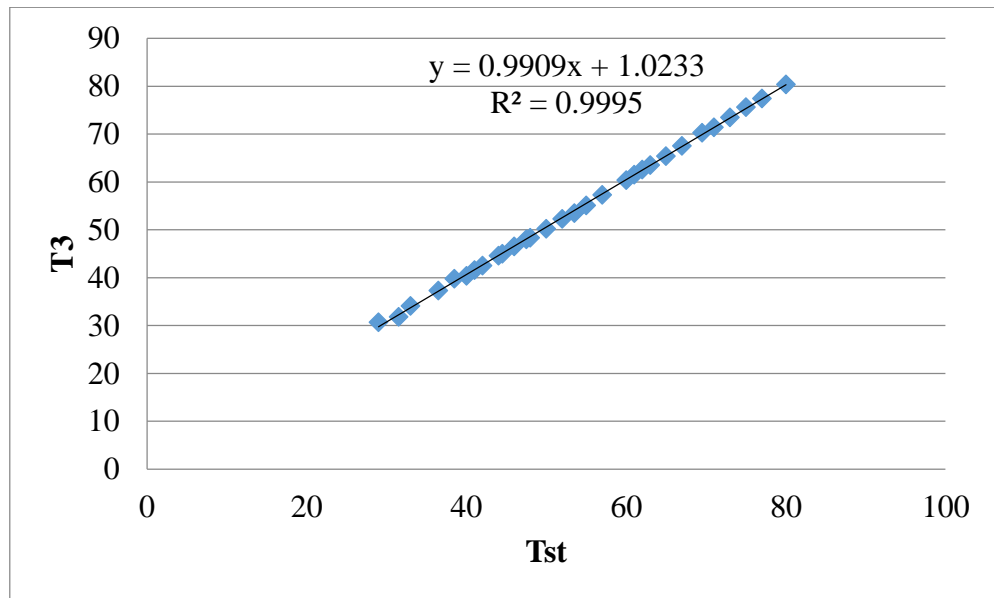
Material	Emissivity
Aluminum foil	0.07
Anodized aluminum	0.82
Polished copper	0.03
Polished gold	0.03
Polished silver	0.02
Polished stainless steel	0.17
Black paint	0.98
White paint	0.90
White paper	0.92–0.97
Asphalt pavement	0.85–0.93
Red brick	0.93–0.96
Human skin	0.95
Wood	0.82–0.92
Soil	0.93–0.96
Water	0.96
Vegetation	0.92–0.96



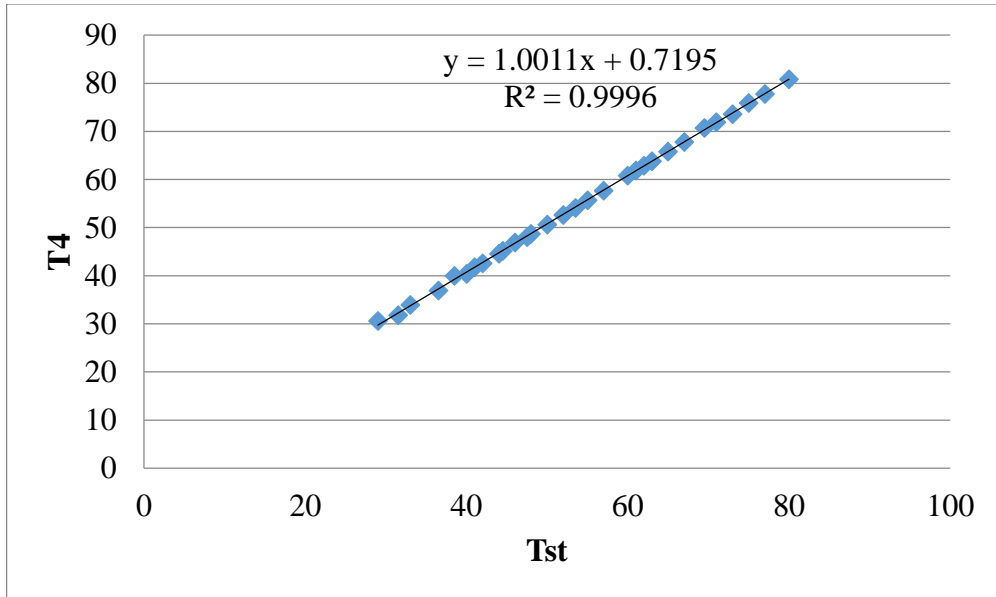
Lampiran 9. Grafik yang terbentuk dari data kalibrasi T<sub>2</sub>



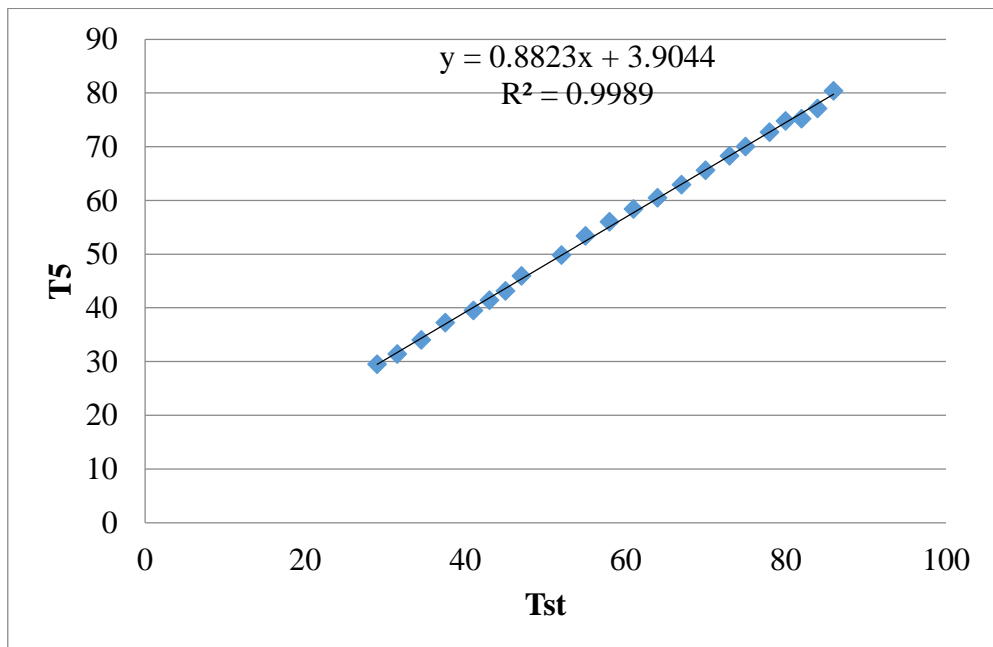
Lampiran 10. Grafik yang terbentuk dari kalibrasi T<sub>3</sub>



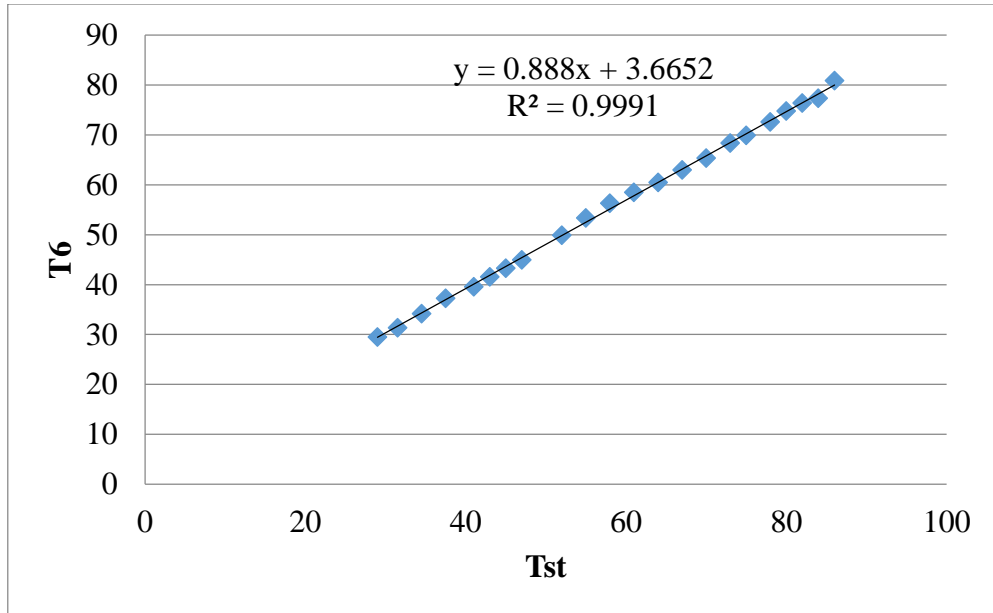
Lampiran 11. Grafik yang terbentuk dari dari kalibrasi T<sub>4</sub>



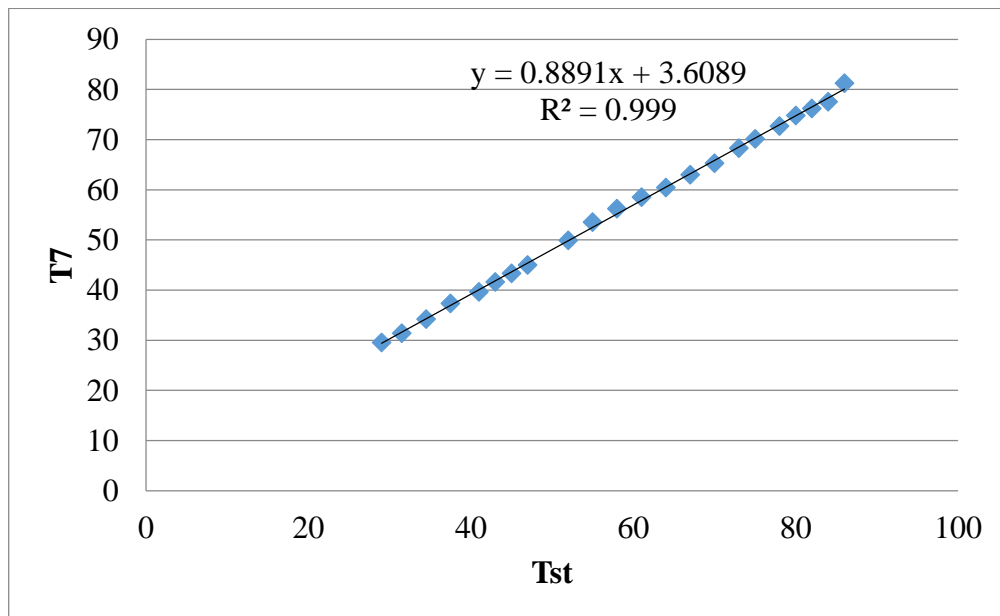
Lampiran 12. Grafik yang terbentuk dari dari kalibrasi T<sub>5</sub>



Lampiran 13. Grafik yang terbentuk dari dari kalibrasi T<sub>6</sub>



Lampiran 13. Grafik yang terbentuk dari dari kalibrasi T<sub>7</sub>



Lampiran 14. Grafik yang terbentuk dari dari kalibrasi T<sub>8</sub>

